

nucleoside residue in the axial position on it. The principal difference, apart from the base, lies in the disposition of the tetrahydropyran ring with respect to the O(2')-C(2'') bond. This amounts to an interchange of the C(3'') and O(1'') atoms and a rotation of 90° about C(2'')→O(2'). We may assume that the tetrahydropyran ring in the more laevorotatory isomer of ThPU will have a similar disposition to that in ThPA(−) so that ThPU(−) would have the molecular structure shown in Fig. 6. A comparison of the pictures of the two isomers (Fig. 6) shows that ThPU(+) can form an internal hydrogen bond [O(3')-H(O3')-O(1'')] with very little distortion of the molecule, the distance O(3')-O(1'') in the crystal being only 3.4 Å. By contrast, in ThPU(−), a large rotation would be necessary to bring O(3') and O(1'') within hydrogen-bonding distance, and such rotation would lead to unacceptably short atomic contact distances, especially between H(12') and H(11') or O(3'). Thus in a non-polar solvent, ThPU(+) is able to present a less polar exterior than ThPU(−) with the resultant observed differences in their physical properties.

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The Crystal Structure of Hexalead Pentaborate, $6\text{PbO} \cdot 5\text{B}_2\text{O}_3$

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Hexalead pentaborate, $6\text{PbO} \cdot 5\text{B}_2\text{O}_3$, is triclinic with the following unit-cell dimensions at 22°C: $a = 6.833 \pm 0.002$, $b = 7.010 \pm 0.002$, $c = 11.125 \pm 0.002$ Å, $\alpha = 85.89 \pm 0.03$, $\beta = 102.00 \pm 0.02$, $\gamma = 119.56 \pm 0.02$. The space group is $P\bar{1}$ with one formula unit in the cell and a calculated density of 6.183 g cm $^{-3}$. Three-dimensional X-ray intensity data were measured with a Picker automatic single-crystal diffractometer and Mo $K\alpha$ radiation. The structure was determined by the heavy-atom method. A full-matrix least-squares refinement resulted in an R value of 0.065. The structure contains isolated $\text{B}_{10}\text{O}_{21}^{12-}$ polyions. These ions are built up from two diborate groups linked by two BO_3 triangles. The lead atoms are covalently bonded to oxygen, as evidenced by the short lead-oxygen distances (2.23 Å and upwards) as well as the asymmetric coordination of the lead atoms.

Introduction

The system lead oxide-boron oxide exhibits two phases of unusual stoichiometry. Geller & Bunting (1937) reported the existence of 1:2, 5:4, 2:1 and 4:1 phases in this system. Bauer & Plötscher (1967) claimed to

have found in addition 1:1 and 9:8 phases. Crystalline compounds with the 5:4 and 9:8 stoichiometric ratios of metal oxide to boron oxide have not been found for other borate systems. The present investigation was undertaken to study the nature of the phase which Bauer & Plötscher believed to be a $9\text{PbO} \cdot 8\text{B}_2\text{O}_3$ com-

Table 1. Final observed and calculated structure factors
The columns are l , $10F_o$ and $10F_c$ respectively.

$-6+L$	6	208	71	0	406	356	$-6+L$	2	756	63	11	498	424	$-6+L$	9	1952	1919	$-6+L$	9	478	297	4	1454	187	9	767	2645	18	360	537								
6	209	22	9	536	466	1	1109	1115	-8	964	549	1	978	1937	-8	1049	1031	-8	2648	2587	-8	634	622	10	393	245	4	1454	187	9	767	2645	18	360	537			
5	151	1503	9	1228	1300	4	1030	1082	-8	15	23	1	1282	1307	-8	432	343	-8	506	513	-8	177	230	-8	205	230	-8	126	184	4	1454	187	9	767	2645	18	360	537
1	516	570	$-8+L$	9	555	403	1	1177	1217	-8	964	549	1	978	1937	-8	1049	1031	-8	2648	2587	-8	634	622	10	393	245	-8	1454	187	9	767	2645	18	360	537		
1	516	570	$-8+L$	9	555	403	1	1177	1217	-8	964	549	1	978	1937	-8	1049	1031	-8	2648	2587	-8	634	622	10	393	245	-8	1454	187	9	767	2645	18	360	537		
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1	516	570	$-8+L$	9	555	403	1	1177	1217	-8	964	549	1	978	1937	-8	1049 </td																					

Table 1 (*cont.*)

pound. The result of our structure determination however, shows the suggested chemical composition to be in error. The compound actually has the formula $6\text{PbO} \cdot 5\text{B}_2\text{O}_3$. This is still a stoichiometry not previously encountered in crystalline binary borates.

Experimental

Hexalead pentaborate, $6\text{PbO} \cdot 5\text{B}_2\text{O}_3$, was obtained by devitrifying at 460–500°C a lead borate glass containing a small amount of lead oxide in excess of the metaborate composition. The glass was prepared as described by Bauer & Plötscher (1967). An X-ray powder diagram indicated the crystalline phase obtained to be identical with the phase believed by Bauer & Plötscher (1967) to be $9\text{PbO} \cdot 8\text{B}_2\text{O}_3$.

A single crystal ground to a sphere of diameter 0.013 cm, was used for the data collection. The maximum deviation from the average diameter was less than 5%. The intensities were measured with a Picker on-line single-crystal automatic diffractometer. Mo $K\alpha$ radiation was used and the X-ray reflexions were measured at 2874 independent reciprocal lattice points.

Unit-cell dimensions and standard errors, $a=6.833 \pm 0.002$, $b=7.010 \pm 0.002$, $c=11.125 \pm 0.002$ Å, $\alpha=85.89^\circ \pm 0.03^\circ$, $\beta=102.00 \pm 0.02^\circ$, $\gamma=119.56 \pm 0.02^\circ$, were ob-

tained by the method of least squares from angle data recorded at 22°C for 12 high-angle reflexions (based on the wavelength 0.7093 Å for Mo $K\alpha_1$). With one formula unit of $6\text{PbO}\cdot 5\text{B}_2\text{O}_3$ the calculated density is 6.183 g cm⁻³. A crude experiment showed the material to have a density of 6.5 ± 0.5 g cm⁻³ in fair agreement with the calculated value.

Structure determination

The observed intensities were converted to structure factors in the usual manner. The linear absorption coefficient is 547 cm^{-1} and a correction for absorption was applied, assuming the crystal to be a perfect sphere. A three-dimensional Patterson map was used to locate the lead atoms and the remaining atoms were clearly seen in an electron-density difference map. The resulting preliminary structure indicated a centre of symmetry.

Refinement of the structure was then carried out with a full-matrix least-squares program from the X-RAY System (Stewart, Kruger, Kundell & Baldwin, 1971). All 2775 independent reflexions observed as larger than the background, were included in the refinement. The atomic scattering factors for O and B (neutral atoms), used for obtaining the calculated structure

Table 2. *Final atomic parameters*

The isotropic and anisotropic temperature-factors are of the form:

$\exp[-U_8\pi^2(\sin\theta/\lambda)^2]$ and $\exp[-2\pi^2[U_{11}(ha^*)^2 + U_{22}(kb^*)^2 + U_{33}(lc^*)^2 + 2U_{12}hka^*b^* + 2U_{13}hla^*c^* + 2U_{23}klb^*c^*]]$

corresponding to the Cruickshank (1956) U notation. Positional parameters are expressed as fractions of the cell edges. All values are multiplied by 10^4 .

	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i> ₁₁	<i>U</i> ₂₂	<i>U</i> ₃₃	<i>U</i> ₁₂	<i>U</i> ₁₃	<i>U</i> ₂₃
Pb(1)	4831 (1)	1405 (1)	3022 (1)	51 (3)	60 (3)	92 (4)	11 (3)	15 (3)	-11 (3)
Pb(2)	-2069 (1)	7618 (1)	4096 (1)	55 (4)	61 (3)	82 (3)	13 (3)	-1 (3)	-21 (3)
Pb(3)	4886 (1)	7326 (1)	615 (1)	73 (4)	64 (3)	89 (3)	21 (3)	41 (3)	-3 (3)

Table 2 (cont.)

	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i>
O(1)	-1679 (31)	3211 (29)	2012 (16)	112 (32)
O(2)	-2780 (31)	-447 (29)	2354 (16)	125 (32)
O(3)	1187 (25)	2324 (23)	2944 (13)	17 (24)
O(4)	5121 (28)	4806 (25)	2888 (14)	64 (27)
O(5)	3257 (30)	5683 (27)	4196 (15)	104 (31)
O(6)	2254 (32)	7804 (30)	5329 (17)	146 (34)
O(7)	321 (27)	6627 (24)	3234 (14)	47 (26)
O(8)	2273 (28)	5786 (26)	1981 (15)	78 (28)
O(9)	-590 (31)	6795 (28)	1080 (16)	112 (31)
O(10)	3042 (29)	9347 (27)	540 (15)	93 (30)
O(11)	0	0	0	361 (79)
B(1)	2992 (45)	4694 (41)	2976 (23)	88 (43)
B(2)	90 (44)	5657 (41)	2026 (23)	80 (42)
B(3)	-1059 (51)	1629 (47)	2417 (26)	141 (50)
B(4)	1873 (39)	6661 (36)	4262 (20)	31 (36)
B(5)	940 (37)	8737 (34)	563 (19)	12 (34)

factors, were those of *International Tables for X-ray Crystallography* (1962). The atomic scattering factors for neutral Pb were obtained from Hanson, Herman, Lea & Skillman (1964) and were corrected for dispersion. The imaginary dispersion terms (Cromer, 1965) were also used in the refinement. The vibrations of boron and oxygen were assumed to be isotropic. A weighting scheme based on the statistical counting errors, compounded with errors assumed to be 1% of the observed intensity, was used. The structure was refined to an *R* value of 0.065 (0.081 for the weighted *R* value), assuming the presence of a centre of symmetry, *i.e.* space group $P\bar{I}$. The ratio of maximum

shift to standard deviation during the last cycle was 0.08. (An attempt was made to refine the structure without the restriction of a centre of symmetry, but the refinement did not converge properly.) Observed and calculated structure factors are given in Table 1. The final atomic coordinates and thermal parameters are given in Table 2.

Discussion of the structure

The anion structure of hexalead pentaborate consists of $B_{10}O_{21}^{12-}$ isolated units (Figs. 1 and 2). This new anion type is the largest isolated borate anion found so far in an anhydrous borate. The anion may be described as two diborate groups connected to each other by a chain section consisting of two BO_3 -triangles. The diborate group was first found by Morimoto (1956) in the mineral borax and has since been found in several borates, both anhydrous and hydrated. Each diborate subunit in the hexalead pentaborate anion is terminated by three oxygen atoms bonded only to one boron, not counting bonds to the lead atoms. Isolated diborate groups have previously only been found with hydroxyl units as terminal atoms. Boron–oxygen bond lengths and bond angles are given Table 3. The values are normal within the error limits. As usual the longest boron–oxygen distances are found for the tetrahedrally coordinated boron atoms.

The lead coordination has noteworthy features. The three crystallographically different lead atoms are all coordinated by three or four oxygens at a short distance (2.23–2.55 Å) and oxygens at longer distances, upwards of 2.76 Å. All the short lead–oxygen distances of a given lead atom fall within the same hemisphere around lead. Moreover the short lead–oxygen distances, with two exceptions [O(8) and O(7)], involve terminal oxygen atoms, *i.e.* oxygens bonded to only one boron. One exception is the lead–oxygen bond to the bridging O(8) atom. This atom, however, is engaged in the two relatively weak boron–oxygen bonds of a pair of BO_4 -tetrahedra.

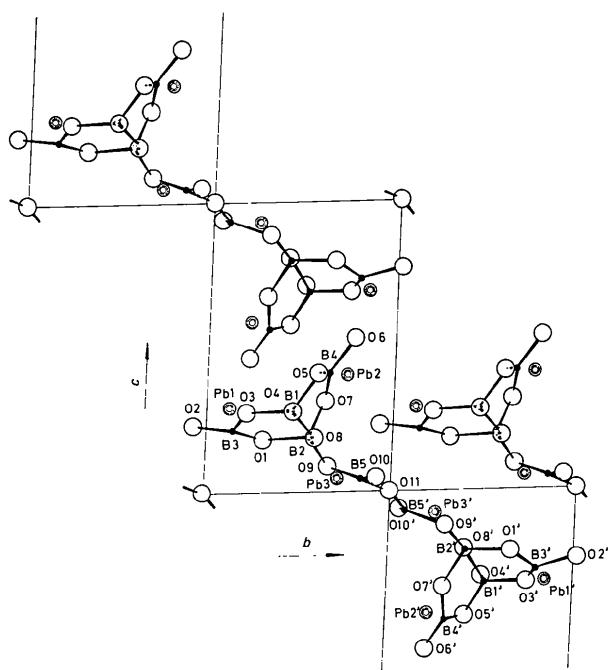


Fig. 1. View of the structure of $6PbO \cdot 5B_2O_3$ along the *a* axis. The concentric circles represent lead, the small filled circles boron and the remaining circles oxygen. The atom numbering corresponds to the numbering used in the text.

Table 3. Interatomic distances and bond angles

B(1)–O(3)	1.50 (3) Å	Pb(1)–O(6)	2.28 (2) Å
B(1)–O(4)	1.44 (4)	Pb(1)–O(4)	2.29 (2)
B(1)–O(5)	1.50 (3)	Pb(1)–O(1)	2.55 (2)
B(1)–O(8)	1.44 (3)	Pb(1)–O(2)	2.76 (3)
B(2)–O(1)	1.53 (3)	Pb(1)–O(3)	2.85 (2)
B(2)–O(7)	1.49 (3)	Pb(1)–O(10)	2.94 (2)
B(2)–O(8)	1.46 (4)	Pb(2)–O(4)	2.23 (1)
B(2)–O(9)	1.41 (3)	Pb(2)–O(2)	2.39 (2)
B(3)–O(1)	1.39 (4)	Pb(2)–O(7)	2.44 (2)
B(3)–O(2)	1.34 (3)	Pb(2)–O(5)	2.79 (2)
B(3)–O(3)	1.36 (3)	Pb(2)–O(5)	2.81 (2)
B(4)–O(5)	1.43 (4)	Pb(2)–O(6)	2.93 (2)
B(4)–O(6)	1.39 (3)	Pb(3)–O(10)	2.30 (2)
B(4)–O(7)	1.38 (3)	Pb(3)–O(2)	2.33 (2)
B(5)–O(9)	1.42 (3)	Pb(3)–O(8)	2.38 (2)
B(5)–O(10)	1.29 (3)	Pb(3)–O(10)	2.48 (2)
B(5)–O(11)	1.38 (3)	Pb(3)–O(4)	3.00 (2)
		Pb(3)–O(1)	3.19 (2)
		Pb(3)–O(9)	3.26 (1)

Table 3 (cont.)

O(3)—B(1)—O(4)	108 (2) ^o	O(2)—Pb(3)—O(10')	83 (1) ^o
O(3)—B(1)—O(5)	107 (2)	O(8)—Pb(3)—O(10')	144 (1)
O(3)—B(1)—O(8)	109 (2)	O(10')—Pb(3)—O(4)	146 (< 1)
O(4)—B(1)—O(5)	109 (2)		
O(4)—B(1)—O(8)	112 (2)	B(2)—O(1)—B(3)	122 (2) ^o
O(5)—B(1)—O(8)	111 (3)	B(1)—O(3)—B(3)	122 (2)
O(7)—B(2)—O(1)	107 (2)	B(1)—O(5)—B(4)	118 (2)
O(7)—B(2)—O(8)	108 (2)	B(4)—O(7)—B(2)	122 (2)
O(7)—B(2)—O(9)	108 (2)	B(1)—O(8)—B(2)	113 (2)
O(8)—B(2)—O(1)	107 (2)	B(2)—O(9)—B(5)	125 (2)
O(8)—B(2)—O(9)	113 (2)		
O(1)—B(2)—O(9)	113 (2)	O(6)—Pb(1)—O(4)	97 (1)
O(1)—B(3)—O(2)	116 (3)	O(6)—Pb(1)—O(1)	78 (1)
O(1)—B(3)—O(3)	118 (2)	O(4)—Pb(1)—O(1)	84 (1)
O(2)—B(3)—O(3)	126 (3)	O(4)—Pb(2)—O(2)	80 (1)
O(5)—B(4)—O(6)	119 (2)	O(4)—Pb(2)—O(7)	82 (< 1)
O(5)—B(4)—O(7)	120 (2)	O(2)—Pb(2)—O(7)	96 (1)
O(6)—B(4)—O(7)	121 (3)	O(10)—Pb(3)—O(2)	83 (1)
O(9)—B(5)—O(10)	126 (2)	O(10)—Pb(3)—O(8)	72 (1)
O(9)—B(5)—O(11)	115 (2)	O(10)—Pb(3)—O(10')	73 (1)
O(10)—B(5)—O(11)	119 (2)	O(2)—Pb(3)—O(8)	81 (1)

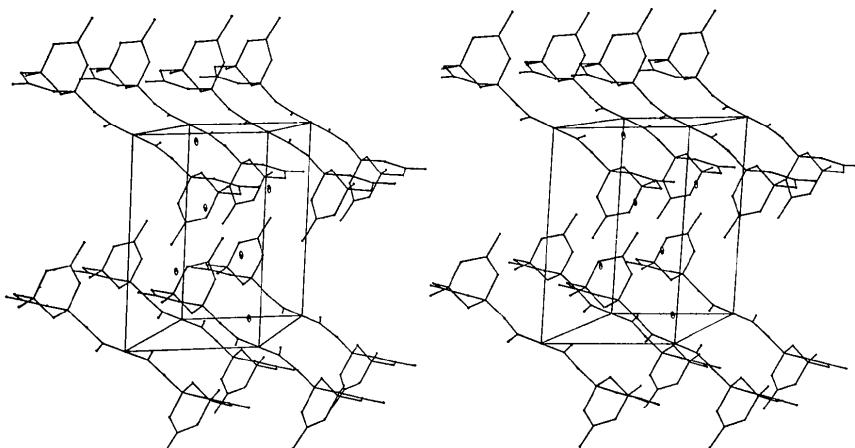


Fig. 2. A stereo pair showing the packing of the structure. The boron–oxygen bonding and the lead atoms are shown.

These characteristics point to the covalent nature of the lead–oxygen bonding. The shortest lead–oxygen distance of 2.23 Å has a value similar to the short lead–oxygen distance of 2.18 Å in lead oxide (Leciejewicz, 1961; Kay, 1961). An ionic lead–oxygen distance, in comparison, would be about 2.60 Å (Pauling, 1960). The asymmetric coordination around the lead atoms in lead oxide, caused by the two 6s electrons, is also retained in the borate compound studied here, emphasizing the similarity with lead oxide in the nature of the bonding.

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